

2.5 Temperature

- The Celsius scale
 - 0°C → freezing point of water
 - 100°C → boiling point of water
 - The Fahrenheit scale
 - 0°F → freezing point of salt/water mixture
 - 100°F → body temperature
 - water freezes at 32°F and boils at 212°F
- ⇒ 100 Celsius degrees ↔ 180 Fahrenheit degrees

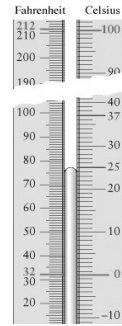


Fig. 2.5

$$\left(\frac{180^\circ\text{F}}{100^\circ\text{C}}\right) = \left(\frac{9^\circ\text{F}}{5^\circ\text{C}}\right)$$

$$T^\circ\text{F} = \left(\frac{9^\circ\text{F}}{5^\circ\text{C}}\right)T^\circ\text{C} + 32^\circ\text{F}$$

$$T^\circ\text{C} = \left(\frac{5^\circ\text{C}}{9^\circ\text{F}}\right)(T^\circ\text{F} - 32^\circ\text{F})$$

- The Kelvin scale - absolute temperature scale
 - 0 K → lowest possible temperature
 - 0 K = -273.15°C
 - same size of the graduation as Celsius
- ⇒ water freezes at 273.15 K and boils at 373.15 K

$$T\text{ K} = T^\circ\text{C} + 273.15$$

$$T^\circ\text{C} = T\text{ K} - 273.15$$

Example:

- Convert -40°F in °C and K.
- $T^\circ\text{C} = (5^\circ\text{C}/9^\circ\text{F}) \times [-40^\circ\text{F} - 32^\circ\text{F}] = (5/9) \times (-72)^\circ\text{C} = -40^\circ\text{C}$
- $T\text{ K} = -40^\circ\text{C} + 273.15 = 233\text{ K}$

2.6 Uncertainty of Measurements

- Represents the reliability of measurements
- Reported as: **number ± uncertainty** (4.88 ± 0.05 kg)
- If not reported: assume **±1 in the last reported digit** (3.7 cm → 3.7 ± 0.1 cm)
- Exact numbers - no uncertainty (5 tables, 10 apples, 1 min = 60 s, 1 in = 2.54 cm)

- Significant figures - digits of a number known with some degree of certainty
 - all non-zero digits
 - all zeros after the first non-zero digit
 - exception - trailing zeros in numbers without decimal point are not significant

Examples:

$$1.32 \rightarrow 3\text{ sf}$$

$$0.005030 \rightarrow 4\text{ sf}$$

$$4500 \rightarrow 2\text{ sf}$$

$$4500. \rightarrow 4\text{ sf}$$

- Scientific notation - representation in the form $\rightarrow A \times 10^a$
 - A - decimal number between 1 and 10
 - a - positive or negative integer
- Examples:
 - $0.00134 = 1.34 \times 10^{-3}$
 - $134 = 1.34 \times 10^2$
 - all digits in A are significant

Table 2.4 Examples of significant figures

Decimal notation	Scientific notation	Number of sf
0.751	7.51×10^{-1}	3
0.007 51	7.51×10^{-3}	3
0.070 51	7.051×10^{-2}	4
0.750 100	$7.501 00 \times 10^{-1}$	6
7.5010	7.5010	5
7501	7.501×10^3	4
7500	7.5×10^3	2*
7500.	7.500×10^3	4

*In this text, treat trailing zeros as insignificant unless they are followed by a decimal point or other information is given. sf, significant figure.

- Significant figures in calculations
 - rounding off (only at the end of a calculation)
 - round up, if next digit is **above 5**
 - round down, if next digit is **below 5**
 - round to the nearest even number, if next digit is **equal to 5** and it is the last digit in the number (if 5 is not the last digit, round up)

Examples: Round to 3 sf.

$3.7643 \rightarrow 3.76$
 $3.7683 \rightarrow 3.77$
 $3.7653 \rightarrow 3.77$
 $3.765 \rightarrow 3.76$

- Addition and subtraction
 - the number of decimal places in the result is the same as the smallest number of decimal places in the data

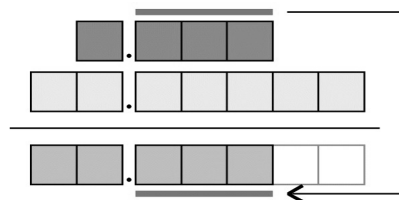


Fig. 2.9

- Multiplication and division
 - the number of significant figures in the result is the same as the smallest number of significant figures in the data

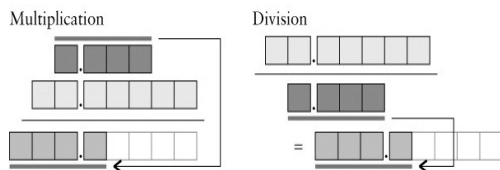


Fig. 2.10

Examples:

$$0.0354 + 12.1 = 12.1 \leftarrow (12.1354)$$

$$5.7 \times 0.0651 = 0.37 \leftarrow (0.37107)$$

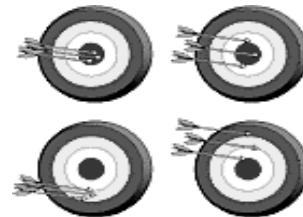
$$5.7 / 0.0651 = 88 \leftarrow (87.55760369)$$

$$3.568 \text{ in} \times (2.54 \text{ cm} / 1 \text{ in}) = 9.063 \text{ cm}$$

2.7 Accuracy and Precision

- Precision - agreement among repeated measurements
 - random error - deviation from the average in a series of repeated measurements
 - small random error $\leftrightarrow$ high precision
 - high precision $\leftrightarrow$ more sf in the result

- Accuracy - agreement of a measurement with the true or accepted value
 - systematic error - deviation of the average from the true value (present in the whole set of measurements)
 - small systematic error $\leftrightarrow$ high accuracy



Example:

- A car is moving at exactly **60 mi/hr**. Compare the precision and accuracy of the following two series of speed measurements using two different radars.
 - A $\rightarrow$ 61.5, 58.3, 62.7, 63.5, 57.1 (average 60.6)
 - B $\rightarrow$ 62.0, 62.5, 61.8, 62.2, 62.1 (average 62.1)
 - A $\rightarrow$ **more accurate, less precise**
 - B $\rightarrow$ **less accurate, more precise**